

CHEMISTRY 2022

Max. Marks: 60

Time: 3 Hour

SECTION "A" MULTIPLE CHOICE QUESTIONS (MCQs)

Marks: 24

NOTE: (i) Attempt all the Questions of this Section.

(ii) Do not copy down the part questions. Write only the answer against the proper number of the Question and its part according to the question Paper.

(iii) Each question carries 1 Mark

1. Choose the correct Answer for each from the given question:

1. Blood is one of the following:

- * Mixture✓ * Compound * Element * Metal

2. Empirical formula of CH_3COOH is:

- * $C_2H_4O_2$ * CH_2O ✓ * CH_4CO_2 * $CHOC$

3. It is considered as the fourth state of matter:

- * Solid * Liquid * Gas * Plasma✓

4. The mass of this particle is almost equal to that of proton:

- * Electron * Neutron✓ * α -Particles * β -particles

5. Line spectrum of Hydrogen atom was studied by:

- * Neil Bohr✓ * Albert Einstein

- * De Broglie * Max Planck

6. This is the isotope of Hydrogen which contains greater number of neutron:

- * Protium * Deuterium * Tritium✓ * Hydrogen

7. The radioactive element in I-A group is:

- * Rubidium * Francium✓ * Cesium * Potassium

8. Dobereiner arranged elements in:

- * Octaves * Triads✓
* Groups and Periods * Sub-Groups

9. All transition elements are:

- * Gases * Metals✓ * Non-Metals * Metalloids

10. According to the modern periodic law presented by Moseley, the properties of elements are the periodic function of:

- * Atomic Size * Atomic Mass
* Atomic Diameter * Atomic Number✓

11. Synthetic resin binds two surfaces together by this bonding:

- * Hydrogen Bonding✓ * Ionic Bonding
* Covalent Bonding * Metallic Bonding

12. Nobel gases are stable because their valance shell contains these numbers of electron: (except Helium)

- * 4 * 5 * 8✓ * 10

13. At room temperature, most of the ionic compounds are:

- * Amorphous Solid * Crystalline Solid✓ * Liquid * Gas

14. 1 atm = _____ Pascal

- * 10325 * 106075 * 10523 * 101325✓

15. Which of the following gas diffuses faster?

- * Oxygen * Chlorine * Fluorine * Helium✓

16. Benzene is an example of this type of solvent:

- * Polar * non polar✓ * ionic * universal

17. 69 g of Sodium (Na) is equal to:

- * 1 moles * 2 moles * 3 moles✓ * 4 moles

18. The size of particles in suspension is:

- * 10^3 nm * 10^2 nm * Less than 10^3 nm * Greater than 10^3 m✓

19. This one of the following is non-electrolyte:

- * Aqueous HCl * Urea✓
* Aqueous NaCl * Molten KCl

20. This one is an alloy:

- * Graphite * Steel✓ * Mercury * Water

21. 1 g equivalent weight of Al is equal to:

- * 1g * 9g✓ * 27g * 54g

22. This gas turns lime water milky

- * CO_2 ✓ * O_2 * NO_2 * N_2

23. This halogen exists in liquid state:

- * Fluorine * Chlorine * Iodine * Bromine✓

24. This element is metalloid:

- * Br * Si✓ * S * Sr